

## Disturbing Equilibrium Lab 38 Answers

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It is possible, however, to disturb that equilibrium by changing conditions. For example, you could increase the concentration of one of the products, or decrease the concentration of one of the reactants, or change the temperature. When a change of this type is made in a reaction at equilibrium, the reaction is no longer in equilibrium.

~~15.7. Disturbing a Reaction at Equilibrium: Le Chatelier's ---~~

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1. State the most important safety concern in this lab and the required precaution you took. 2. Write a balanced net ionic equation for the equilibrium that existed when iron(III) ions and thiocyanate ions were combined in the beaker. Note the color of each reactant and product. 3. For each of the test tubes 2-5:

~~LE CHATELIER'S PRINCIPLE EXPERIMENT 26~~

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Part 3: Aqueous Ammonia Solution. Instructor Prep: At the beginning of lab prepare a stock solution of aqueous ammonia. Add 4 drops of concentrated 15 M  $\text{NH}_3$  (aq) and 3 drops of phenolphthalein to a 150-mL (medium) beaker, top it up with 100-mL of distilled water, and mix with a stirring rod. Label the beaker and place it on the front desk.

~~12. Equilibrium and Le Chatelier's Principle (Experiment ---)~~

Likewise, if you change the conditions of an equilibrium, it will change in a way that partially undoes whatever it is you did to it in the first place. Change in Concentration Let's say that you're doing a reaction with the equation  $A + B \rightleftharpoons C$ . Le Chatelier's principle says that if we change the concentration, the position of the equilibrium will also change.

~~Chemistry: Disturbing Equilibrium: Le Chatelier's Principle~~

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