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- 3 9. Explain the difference between a bonding electron domain and a nonbonding electron domain using the examples in Model 1.
10. Circle the correct word or phrase to

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complete the
sentences: Pairs of
electrons will
(attract/repel ...

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shapes are tetrahedral, trigonal planar, trigonal pyramidal, bent, and linear and are named by measuring the bond angles between the central atom and another atom bonded to it.

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In this activity you
will learn how to
predict molecular

shapes. Model 1 –

Lewis Structures Lone

pair = •• H₂CO₃

electron domains (3

bonding, 0

nonbonding) 3-D

Molecular Shape BeF₂

2 electron domains

(2 bonding, 0

nonbonding) CH₄

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electron domains (4
bonding, 0
nonbonding) NH_3 4

electron domains (3
bonding, 1

nonbonding) H_2O 4
electron domains

~~20 Molecular~~

~~Geometry S Mrs.~~

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~~Classes~~

electron pair

geometry: molecular

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geometry: shape:
polarity: BeH₂: 2: 2: 0:
linear: MX₂: linear:
nonpolar: BH₃: 3: 3: 0:
trigonal planar: MX₃:
trigonal planar:
nonpolar: SO₂: 2: 2: 1:
trigonal planar: MX₂
E: bent: polar : CH₄:
4: 0: tetrahedral: MX₄:
4: tetrahedral:
nonpolar: NH₃: 3: 3: 1:
tetrahedral : MX₃ E:
trigonal . pyramidal.

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polar: H_2O : 2: 2:

tetrahedral: MX_2E_2 :

bent or . V-shaped.

polar: PCl_5 : 5: 0:

trigonal.

bipyramidal. MX_5 :

trigonal .

bipyramidal.

nonpolar: SF_6 : 6: 0

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How to Determine

Molecular Geometry

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– YouTube: This video describes one method for quickly finding the major geometrical shapes for simple molecules.

Molecular Geometries The VSEPR theory describes five main shapes of simple molecules: linear, trigonal planar, tetrahedral, trigonal

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bipyramidal, and
octahedral.

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