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Overview: Philippians *CH 13*

CHEMISTRY KINETIC MOLECULAR

THEORY Chemistry: Charles's Law

(Gas Laws) with 2 examples |

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~~Nurses and Nursing Students Be Lazy! Don't Memorize the Gas Laws!~~ *Chapter 13 Study Guide Gases*

Chapter 13 - Gases 195 Exercise 13.3 - Equation Stoichiometry: Iron is combined with carbon in a series of reactions to form pig iron, which is about 4.3% carbon.

$$2C + O_2 \rightarrow 2CO$$
$$Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$$

2CO + C (in iron) \rightarrow CO₂ Pig iron is easier to shape than pure iron, and the presence of carbon lowers its melting point

Chapter 13 - Gases

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Students should be able to... ·

Explain the kinetic theory of matter as it applies to gases. ·

Identify the properties of an ideal gas vs. a real gas. · Know units of

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pressure in atm, torr, mm Hg, kPa, psi, and in Hg. · Calculate using Boyle's Law, Charles' Law, Gay-Lussac's Law, Combined Gas Law, and Ideal Gas Law (using 0.0821 for R), and Graham's Law.

Chapter 13 Study Guide: Gases - Tripod

194 Study Guide for An Introduction to Chemistry Section Goals and Introductions Section 13.1 Gases and Their Properties Goals • To describe the particle nature of both real and ideal gases. • To describe the properties of gases that can be used to explain their characteristics: volume, number of particles, temperature, and pressure.

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A gas can flow into a space
occupied by another gas. 9. The
diffusion of a gas is caused by the
random motion of the particles of
the gas. 10. Lighter gas particles
diffuse less rapidly than do
heavier gas particles. 11. During

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effusion. a gas escapes through a tiny opening into a vacuum. 12.

13 STUDY GUIDE FOR CONTENT MASTERY 13 STUDY GUIDE FOR ...

Chapter 13 Study Guide Answers

1. Describe the assumptions/postulates of the kinetic-molecular theory of gases:

- Gases are composed of tiny particles in constant rapid, (“random” or “straight-lined”?) motion.
- Gases are separated by relatively huge distances. The volume of the particles is essentially zero.
-

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read about gas pressure. Circle
the letter of the choice that best
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13. Pressure is defined as force
per unit a. area. b. mass. c. time.
d. volume. 14.

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